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1. 88% The actual
yield of a reaction is
22 g and the

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theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N_2 are mixed with 12.0 mol of H_2 according to the following equation: $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$...

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45 mol The following
equation represents a
laboratory

preparation for

oxygen gas: 2KClO_3

$3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$

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How many moles of O₂ form if 30 mol of

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Calculate the molar mass of the following, and then write the molar mass in ratio form. KNO_3 .

Na_2SO_4 . $\text{Ca}(\text{OH})_2$

$(\text{NH}_4)_2\text{SO}_3$. CuSO_4 .

Moles and Mass.

Determine the ...

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PLUS. Terms in this

set (36) a balanced

chemical equation

allows one to the

determine the. mole

ratio of any two

substances in the

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reaction. The coefficients in a chemical equation represent the relative numbers of moles of reactants and products. actual ...

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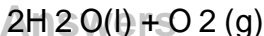
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Answers will vary. An example is included below: $2\text{H}_2\text{O}_2(\text{aq})$



This describes the decomposition reaction of hydrogen peroxide.

Microscopic: Two molecules of hydrogen peroxide (in aqueous solution) decompose to

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produce two molecules of liquid water and one molecule of oxygen gas. Macroscopic: Two moles of hydrogen peroxide (present in ...

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g mgh-E K U g K

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and the theoretical
yield is 25 g Calculate
the percentage yield
2 60 mol of N_2 are
mixed with 120 mol
of H Read Online
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reactants are
converted into ...

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